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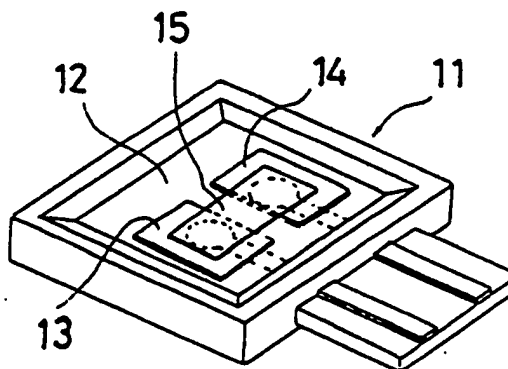
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Method for measuring the ion-concentration and the like, a calibrating sheet and a calibrating method for the measuring device used therefor.

Described is a method for measuring the ion concentration and the like of a sample by impregnating a liquid-absorbing material with the sample and bringing said impregnated material into contact with the measuring portion of a device for measuring the ion concentration and the like. Furtheron, a sheet is disclosed for the calibration of the device for measuring the ion concentration and the like and a method for calibrating the device for measuring the ion concentration and the like measuring device using said sheet.

Fig. 1



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**METHOD FOR MEASURING THE ION-CONCENTRATION AND THE LIKE,  
A CALIBRATING SHEET AND A CALIBRATING METHOD  
FOR THE MEASURING DEVICE USED THEREFOR**

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1 **BACKGROUND OF THE INVENTION**

**Field of the Invention**

The present invention relates to a method of measuring the ion-concentration and the like, a calibrating sheet and a calibrating method for the  
5 measuring device used therefor.

**Description of the Prior Art**

For measuring the ion-concentration, the quantity of dissolved oxygen, the conductivity and the like (hereinafter, ion concentration and the like) in an infinitesimal quantity of a liquid secreted from a living body, different  
10 sampling jigs must be used for every object to be measured to obtain a sample for the measurement.

For example, in order to sample saliva, a cotton roll has been put into the

1 oral cavity to impregnate the cotton roll with the saliva whereafter the cotton  
roll has been subjected to a centrifugal separation by means of a centrifugal  
separator to obtain the pure saliva. To sample the saliva from glandula parotis,  
a cup has been put adherently on the mucous membrane of the oral cavity to  
5 obtain the saliva (prior art 1).

In order to sample sweat emanating from the surface of a body, a quanti-  
tative test method (QPIT method) has been used. For example, sweat perspired  
from the arm has been absorbed by a filter paper and the sweat absorbed in this  
filter paper has been diluted to measure the ion-concentration in the sweat, or,  
10 an ion electrode and conductivity electrode have been directly pressed against  
the skin perspiring sweat (prior art 2).

In addition, to measure the ion-concentration of the skin, in particular  
the pH, a glass electrode having a flat surface, has been pressed against the skin  
(prior art 3).

15 Furthermore, when the deterioration test of iron and steel subjected to a  
pretreatment for painting, painting of cars, painting of buildings, plated artic-  
les and the like is carried out, for example by the measurement of pH, pNa, pCl,  
conductivity and the like of contaminating substances adhering to the painted  
film on iron and steel, contaminating substances have been sampled by mopp-  
20 ing off said painted film in a predetermined area in a predetermined direction  
and a number of times by means of a gauze and by the use of a clean glove, which  
was boiled in distilled water and impregnated therewith, said gauze has been  
boiled in distilled water for 1 hour, and then a total volume 500 ml has been  
made to obtain a predetermined sample (prior art 4).

25 However, the above described prior art methods have the following dis-  
advantages.

In order to obtain a sample to be measured according to the above de-  
scribed prior arts 1, 2 and 4, a great expenditure of instruments is required and  
a pretreatment is necessary. Therefore, the disadvantage occurs in that much  
30 time and labour are required in the measurement. In addition, in the case of the  
prior art 3, the measurement can only be carried out for a portion of the skin  
which can be touched directly and the quantity of KCl leaked is increased,  
whereby the measured result can not be distinguished from the influence of the  
solution.

35 In addition, since according to objects to be measured only a small  
amount of the sample to be measured can be sampled, it is difficult to carry out  
an accurate measurement.

For example, for the calibration of an electrode of a pH meter five kinds of  
standard buffer solutions as calibrating solutions having a pH of 1.68, 4.01,  
40 6.86, 9.18 and 10.02, respectively have been used which have been prepared

- 1 according to JIS at a temperature of 25°C. The above described five kinds of  
standard buffer solutions include an oxalate standard solution (0.05 M aqueous  
solution of potassium tetroxalate  $\text{KH}_3(\text{C}_2\text{O}_4)_2 \cdot 2\text{H}_2\text{O}$ ), a phthalate standard  
solution (0.05 M aqueous solution of potassium hydrogen phthalate  
5  $\text{C}_6\text{H}_4(\text{COOK})(\text{COOH})$ ), a neutral phosphate standard solution (0.025 M aqueous  
solution of mono potassium phosphate  $\text{KH}_2\text{PO}_4$  + 0.025 M aqueous solution of  
disodium phosphate  $\text{Na}_2\text{HPO}_4$ ), a borate standard solution (0.01 M aqueous  
solution of sodium borate (borax)  $\text{Na}_2\text{B}_4\text{O}_7 \cdot 10\text{H}_2\text{O}$ ) and a carbonate standard  
solution (0.025 M aqueous solution of sodium hydrogen carbonate  $\text{NaHCO}_3$  +  
10 0.025 M aqueous solution of sodium carbonate  $\text{Na}_2\text{CO}_3$ ).

Pure water having a conductivity of  $2 \times 10^{-6} \Omega^{-1} \text{cm}^{-1}$  (at 25°C) or less has  
been used for the preparation of the standard buffer solutions. For example, the  
oxalate standard solution is prepared by dissolving 12.71 g of potassium tetra-  
oxalate (JIS K8474, suitable for the use in the measurement of pH) in said pure  
15 water so as to obtain 1 liter in total.

Fig. 15 shows the procedure of the conventional calibrating method.  
"Preservation" means that the prepared standard solution is preserved in a  
bottle made of high quality glass or polyethylene which is tightly closed. A  
calibrating solution including the standard solutions, which are prepared and  
20 preserved in the above described manner is on the market in the form of a set as  
an accessory of a pH meter.

Since only a calibrated pH meter can achieve measurements with the best  
accuracy, it is necessary to carry out the calibration if needed and accordingly,  
the above described calibrating solution is indispensable. However, when the  
25 measurement has to be carried out outside of the laboratory, for example, on  
fields, it is necessary to transport the above described set of calibrating solu-  
tions and washings together with the pH meter. Accordingly, there are the  
disadvantages that the apparatus is remarkably bulky, so that the calibrating  
equipment is remarkably difficult to transport, and that the calibration is  
30 difficult to accurately carry out. In short, the measurement of pH and cali-  
bration have been difficult to carry out with accuracy.

#### SUMMARY OF THE INVENTION

It is an object of the present invention to provide a method of measuring  
35 the ion-concentration and the like with high accuracy, by measuring only a  
small sample and obtaining an accurate measurement without using instru-  
ments for sampling and subjecting to a complicated pretreatment.

It is another object of the present invention to provide a sheet for the use  
in the calibration which is handy and easy to use and a calibrating method for  
40 easily calibrating measuring units of the measuring devices.

1       The first object has been solved by a method for measuring the ion con-  
centration and the like of a sample, which is characterized by impregnating a  
liquid-absorbing material with the sample and bringing said impregnated  
material into contact with the measuring portion of a device for measuring the  
5       ion concentration and the like.

      According to the method of measuring the ion-concentration and the like  
of the present invention, the sample is absorbed in the liquid-absorbing  
material and said absorbing material is brought into contact with the measur-  
ing unit to measure the ion-concentration and the like in said sample, so that a  
10       highly accurate measurement can be obtained, even though the quantity of the  
sample is very small. In addition, the measurement can be surely carried out  
without using instruments for sampling and subjecting to complicated pre-  
treatments.

      The subclaims relate to preferred embodiments of the method of the  
15       invention.

      The other object has been solved by a sheet for the calibration of the de-  
vice for measuring the ion concentration and the like for use in the method  
which is characterized by a chemically stabilized hygroscopic water-absorbing  
sheet eluting no ion impregnated with a calibrating buffer solution and there-  
20       after dried.

      Since the sheet for the use in the calibration of the ion-concentration and  
the like measuring device according to the present invention is obtained by ab-  
sorbing the calibrating buffer solution in the water-absorbing sheet and then  
drying the water-absorbing sheet, the sheet for the use in the calibration of the  
25       ion-concentration and the like measuring device can be preserved in the form  
of a solid sheet. Accordingly, the apparatus is not bulky but extremely handy.  
In addition, the sheet is of sheet-like shape so that it is easy to handle.

      Furtheron, a method for calibrating the device for measuring the ion  
concentration and the like is provided which is characterized in that the sheet  
30       is brought into contact with an electrode of the device for measuring the ion  
concentration and the like and a predetermined quantity of pure water is added  
dropwise to said sheet.

      According to the method of calibrating an ion-concentration and the like  
measuring device according to the present invention, it is only necessary to  
35       bring said sheet for the use in the calibration into contact with an electrode of  
the device for measuring the ion-concentration and the like and to add a pre-  
determined quantity of pure water drop by drop to the sheet, so that a simple  
and secure cali-bration can be obtained.

#### BRIEF DESCRIPTION OF THE DRAWINGS

40       The preferred embodiments of the present invention are shown in the

1 drawings.

Fig. 1 shows a perspective view of one example of an ion-concentration measuring composite electrode for the use in the practice of a method according to the present invention;

5 Fig. 2(A), (B) and (C) show the way how the ion-concentration of saliva is measured;

Fig. 3(A), (B), (C) and (D) show the way how the ion-concentration of tears is measured;

10 Fig. 4(A), (B), (C), (D) and (E) show the way how the ion-concentration of vagina liquid and sperma is measured;

Fig. 5(A) and (B) show the way how the ion-concentration of sweat is measured;

Fig. 6(A) and (B) show the way how the ion-concentration of rain is measured;

15 Fig. 7(A), (B) and (C) show the way how the ion-concentration of skin and hair is measured;

Fig. 8(A), (B) and (C) show the way how the ion-concentration of fish is measured;

20 Fig. 9(A) and (B) show the way how electrolytes in floating substances within a clean room are measured;

Fig. 10(A), (B), (C) and (D) show the way how electrolytes in soil, ash and other powdery samples are measured; and

Fig. 11 shows the way how the ion-concentration of a cloth is measured.

25 Fig. 12 shows a chart of the individual steps from the manufacture of a sheet for the use in the calibration according to the present invention to the calibration by the use of said sheet for the use in the calibration;

Fig. 13 shows a perspective view of one example of said sheet for the use in the calibration; and

30 Fig. 14 shows a diagram describing a method of calibrating an ion-concentration and the like measuring device and

Fig. 15 shows a chart describing the prior art.

#### DESCRIPTION OF THE PREFERRED EMBODIMENTS

35 The preferred embodiments of the present invention will be below described with reference to the drawings.

Fig. 1 shows an example of a sheet type ion-concentration measuring composite electrode and a material easy to absorb a solution used in the practice of a method according to the present invention wherein reference numeral  
40 11 designates a sheet type ion-concentration measuring composite electrode of

1 which a measuring portion 12 is provided with a selective ion-response membrane as a part of measuring electrode 13 and a liquid junction membrane formed of inorganic sintered porous materials or organic high molecular porous materials impregnated with KCl as a part of reference electrode 14.

5 A sheet type ion-concentration measuring composite electrode constructed in this manner is subject of the European patent application 87 113 266.8 filed by the applicant (Horiba Ltd.) on November 27, 1987.

In addition, referring again to Fig. 1, reference numeral 15 designates a liquid-adsorbing material placed as to extend over said measuring electrode 13 and said reference electrode 14. Said material 15 absorbs easily a solution, is chemically stabilized and flexible, and does not influence the measurement, i.e. it emanates no interferential ingredient against an object to be measured. The material 15 is for example selected from materials such as a sponge, absorbent cotton, neutral paper, porous glass, glass fiber and the like and can be used in various shapes, as it will be shown by the following preferred embodiments.

Figs. 2 to 6 show a measuring method wherein liquid samples are used.

Fig. 2 shows the case where the ion-concentration in saliva is measured. As shown in Fig. 2(A), an assembly of an area of 15 mm x 15 mm comprising a sheet type absorbing material 21 and an evaporation-inhibiting high molecular sheet 22 in piles is adhered to a portion of each gland perspiring saliva about ten and several seconds by holding it with tweezers 23 to absorb saliva in said absorbing material 21 (refer to Fig. 2(B)). And then, the assembly is placed on the measuring portion 12 so that one side of the absorbing material 21 may be brought into contact with one side of the electrode 13 (14) to measure pH, pNa, pK, pCl, conductivity and the like of the saliva (refer to Fig. 2(C)).

In addition, reference numeral 24 designates a sealing portion to be sealed by a heat or an ultrasonic wave treatment.

Thus, a sample of saliva of several microliters to several ten microliters is sufficient in quantity to be measured. Advantages occur in that instruments for obtaining the sample are hardly required and a pretreatment is not required as necessary in the prior art 1.

Secretions from a living body, a particularly small amount of tears, mucus, sperm, vagina liquid and the like are preferably sampled with a thin high molecular nonwoven fabric of about several microns to 100 microns as the absorbing material over a wide area of the absorbing material.

In Fig. 3 the sampling of tears (including tear 31 when an eye lotion is added drop by drop and tear 33 on a contact lens (Fig. 3(A)) is shown since only a piece of the thin film nonwoven sheet 34 is difficult to handle after the sampling of the tear, a frame member 35 is provided on the back side (Fig. 3(B)), or, the sheet 34 is provided with a multi-layer construction comprising a

- 1 hydrophilic absorbing sheet 34a and a hydrophobic absorbing sheet 34b to prevent the evaporation from the hydrophobic surface and to improve the adhesion to the electrode 13 (14).

- 5 In addition, reference numeral 36 designates a sealing portion to be sealed by a heat or an ultrasonic wave treatment.

Fig. 4 shows the way how to sample vagina liquid and sperma to measure the pH within the vagina and of sperma.

- 10 As shown in Fig. 4(A), in order to measure pH in the circumference 41 of the womb and the circumference 42 of the vagina which is effective for the discrimination of sexes of an embryo, a fingerstall 43 formed of an absorbing material put on a clean finger is inserted into the vagina to adhere vagina liquid to the external circumference of said fingerstall 43.

- 15 In addition, Fig. 4(B) shows a sack 45 formed of an absorbing material for sampling sperma from penis 44. Figs. 4(C), (D) and (E) show an absorbing material suitable for sampling from merely the circumference of the womb, respectively, Fig. 4(C) shows a sack 46 formed of an absorbing material as a whole, Fig. 4(D) shows a sack 47, of which the end portion is formed of an absorbing material 47a and of which the base end portion is formed of an evaporation-inhibiting high molecular sheet 47b, and Fig. 4(E) shows a sack 48  
20 formed of, for example a rubber and provided with an absorbing material 49 at a part of the point end portion thereof.

In addition, reference numeral 50 designates a sealing portion to be sealed by a heat or an ultrasonic wave treatment.

- 25 Fig. 5 shows the way how to sample sweat. In Fig. 5(A), an evaporation-inhibiting high molecular sheet 51 is provided with a thin film sheet type absorbing material 52 on one surface thereof and with an adhesive layer 53. This absorbing material 52 is adhered to an area where a large amount of sweat is perspired such as the forehead (edge of the hairs) and under the arm, as shown in Fig. 5(B).

- 30 The portion, to which the absorbing material 52 is adhered, is preferably pre-cleaned by for example washing. In addition, in order to perspire sweat, a certain movement is advisable or the temperature within the room should be enhanced. Thus, the perspiration of sweat is promoted to obtain the predetermined amount of sweat in a short time.

- 35 According to the above described method, a small quantity of sweat can be measured and advantages occur in that instruments for obtaining the sample are hardly required and a pretreatment is not required as necessary in the prior art 2.

- 40 In the above described preferred embodiments liquids perspired from a living body, such as human body, are sampled. It is also necessary to measure



1 the acidity of rain water and the like, as shown in Fig. 6.

As shown in Fig. 6(A), a film 62 is stuck to a back surface of a thin film of a nonwoven sheet 61 being several ten microns thick, and a fixing frame 63 which is the same as the fixing frame 35 as shown in Fig. 3(B) is provided, followed by an adhesive layer 64 provided thereon. This thin film of the nonwoven sheet 61 is stuck to the upper surface of a leaf 65 and taken off from the leaf 65 after a certain time from the beginning of raining 66. Then, the thin film of the nonwoven sheet 61 is placed in the electrode 13 (14) to measure pH and the like.

10 In the above described preferred embodiments the samples are liquid, but, also relatively dry samples, such as skin and hair of humans, can be measured.

In Fig. 7(A), reference numeral 71 designates an absorbing material pre-impregnated with a predetermined amount of distilled water or a 0.1 %-NaOH solution which is adjusted to a pH to about 5.0 with HCl (pH of the solution giving a skin a minimum swelling is about 5.0). The resulting absorbing material 71 is sufficiently brought into contact with a predetermined portion of the arm ((Fig. 7(B)) or hair (Fig. 7(C)) several times and then placed on the electrode 13 (14) to carry out the measurement.

20 According to the present invention, the ion-concentration and the like in the skin and hair having a lower water-content can be simply measured in contrast to the prior art 3. In addition, it goes without saying that the liquid which does not influence the measurement is not limited to the above described distilled water and the like.

25 A type of sampling by pre-impregnating an absorbing material with a solution which does not influence the measurement, as in the above described preferred embodiment, can be applied also to the deterioration test of iron and steel subjected to a pretreatment for the painting, painting of cars, painting of buildings, plated articles and the like.

30 Although it is not shown, if a thin film of an absorbing sheet impregnated with a predetermined quantity of distilled water is brought into contact with a surface of painted articles, plated articles and the like to achieve a predetermined sampling and then placed on the electrode 13 (14), pH, pNa, pCl, conductivity and the like can be measured at the same time.

35 In this case, an infinitesimal quantity of distilled water as the non-interferential liquid of about several microliters to 100 microliters is sufficient to be measured. Advantages occur in that the dilution of a sample with a great quantity of water is not required and the application of heat and the like to substances stuck to the surface is not required, whereby the highly accurate measurement can be simply carried out without requiring any other treatments and the measurements over large items can be achieved at the same time.

40

1 A method according to the present invention for measuring electrolytes and the like in foodstuffs, such as fish, beef, pork and chicken is also provided.

Fig. 8 shows a method of measuring the ion-concentration of fish wherein a comparatively dry portion, such as the surface, is measured as shown in Fig. 5 8(A). A sheet 81 formed of an absorbing material impregnated with distilled water is placed on the electrode 13(14) and a sample 82 composed of predetermined portion of a fish is placed on said sheet 81. On the other hand, the measurement of a comparatively wet portion is preferable in that a sample 84 is put in a bag-like absorbing sheet 83 (Fig. 8(B)) and the absorbing sheet 83 is 10 subjected to the absorption on the electrode 13 (14) (Fig. 8(C)).

Fig. 9 shows the case where electrolytes in floating substances within a clean room are measured wherein a filter member (such as filter paper, sponge and porous material) 91 formed of a material eluting no electrolyte is impregnated with distilled water and put in a gas passage 92 for a predetermined time 15 and then placed on the electrode 13(14) for measurement.

According to this method, pNa, pCl and the like in floating substances within a clean room can be simply measured.

In addition, a mesh of the above described filter member 91 is suitably selected according to a flow rate of gas but only a first layer may be composed of 20 a filter 93 formed of an absorbing material and the rest may be composed of the conventional gas filter 94 to form a composite filter, as shown in Fig. 9(B).

Fig. 10 shows a method of measuring electrolytes in soil, ash and other powdery samples wherein a sample, such as soil 101, is poured into a bag 102 formed of an absorbing material, as shown in Fig. 10(A), and then the mouth 25 portion of the bag 102 is sealed (Fig. 10(B)). Subsequently, the bag 102 including the sample 101 is placed on the electrode 13 (14) and distilled water 104 is brought dropwise on the bag 102 (Fig. 10(C)). Alternatively, a sheet-like absorbing material 103 is placed on the electrode 13 (14) and the sample 101 is placed on the sheet-like absorbing material 103, and then distilled water 104 is 30 added dropwise to the sample 101 (Fig. 10(D)).

In each of the above described preferred embodiments the absorbing material is composed of a chemically stabilized material eluting only a small amount of the electrolyte within the error of measurement even though it is immersed in distilled water or a solution. However, when using as a sample 35 substances eluting electrolytes to some extent when distilled water is added dropwise, such as cloths, acidic papers and ceramics, the sample 111 (for example a part of cloths) is directly placed on the electrode 13 (14) and distilled water is added dropwise on the sample 111 for the measurement, as shown in Fig. 11.

40 The preferred embodiments of a calibrating sheet as well as a calibrating

- 1 method of measuring a device for ion-concentration and the like will be below described with reference to Figs. 12 to 14.

Fig. 12 shows the individual steps from the manufacture of the sheet for the use in the calibration according to the present invention to the calibration  
5 by the use of said sheet. The sheets for the use in the calibration of pH 4 and pH 7 will be illustrated.

At first, 10.21 g potassium hydrogen phthalate (JIS K8809, for the use in the measurement of pH), 3.40 g monopotassium phosphate (JIS K9007, for the use in the measurement of pH) and 3.55 g disodium phosphate (JIS K9020, for the use in the measurement of pH) are weighed (refer to S1).  
10

Subsequently, the 10.21 g potassium hydrogen phthalate are dissolved in pure water (hereinafter referred to as pure water for the use in the preparation) having a conductivity of  $2 \times 10^{-6} \Omega^{-1} \text{cm}^{-1}$  (at 25°C) or less to prepare the phthalate standard solution of 1 liter. On the other hand, the 3.40 g mono-  
15 potassium phosphate and the 3.55 g disodium phosphate are dissolved in pure water for the use in the preparation to prepare the neutral phosphate standard solution of 1 liter (refer to S2).

Then, the above described phthalate standard solution and the above described neutral phosphate standard solution are weighed, respectively, and  
20 put into a separate vessel in a predetermined quantity (refer to S3).

Subsequently, the absorbing sheet is immersed in the standard solution within the above described vessel in turn to absorb the solution in the absorbing sheet (refer to S4).

Chemically stabilized hygroscopic materials which do not elute ions, are preferably used as materials of the absorbing sheet. For example, high molecular materials having a good flexibility and a good adhesion to the electrode in addition to the above described properties and inorganic materials, glass fiber and ceramic fiber, may be used.  
25

In the present preferred embodiment, a nonwoven fabric (of about 220 microns thick) formed of cellulose at the rate of 100% was used as the absorbing sheet.  
30

Subsequently, the absorbing sheet impregnated with the standard solution in the above described manner is dried at room temperature to obtain a sheet for use in the calibration 1, 2 of pH 4 and pH 7 (refer to S5 and Fig. 13).

35 In this time, it is important to dry so that dried ingredients of fine particle powders may not be torn off from the absorbing sheet after the evaporation of water.

The sheets of the use in the calibration 1, 2, which were obtained in the above described manner, are suitably packaged and preserved (refer to S6). In  
40 addition, these sheets for the use in the calibration 1, 2 are on the market as an

1 accessory of the pH meter in the form of a set together with sheets for the use in  
the calibration impregnated with other standard solutions.

And, in order to calibrate an electrode of the pH meter in the measure-  
ment of pH, as shown in Fig. 14, the above described sheet for the use in the cali-  
5 bration 1 (2) is placed very tightly on a surface of a pH electrode 3 of flat surface  
construction as shown in for example Japanese Patent Application No. Sho.  
61-285371 by the present applicant (Horiba Ltd.) and pure water of a quantity  
corresponding to a quantity of pure water for the use in the preparation used in  
the preparation of the standard solution by adding pure water as shown by S2  
10 in the above described Fig. 12 is dropped on the above described sheet for the use  
in the calibration 1 (2). This drop by drop addition of the pure water 4 leads to  
the reformation of the standard solution in the same way as it is the case before  
the dehydration in the sheet for the use in the calibration 1 (2) and the cali-  
bration of the pH electrode can be carried out by the use of this standard  
15 solution (refer to S7).

Said pH electrode 3 of flat surface construction was calibrated by the use  
of the sheet for the use the in the calibration obtained by impregnating the  
absorbing sheet made of a nonwoven fabric formed of cellulose at the rate of  
100% used in the present preferred embodiment with the above described stand-  
20 ard solution with good reproducibility and good reproducibility after leaving  
unattended for several days.

As to pH standard buffer solutions, the borate standard solution and the  
carbonate standard solution absorb carbon dioxide gas resulting in a remark-  
ably lowering of pH so that in the case where the absorbing sheet is impregnated  
25 with these standard solutions to obtain the sheet for use in the calibration, it is  
preferable that a process of impregnating the absorbing sheet with the standard  
solution and the following processes are conducted in an atmosphere of inert  
gas and the dehydration as well as the preservation are conducted so as to avoid  
the contact with air.

30 In addition, when the sheet for the use in the calibration is prepared  
without using the pH standard buffer solution according to JIS, it is preferable  
that the absorbing sheet is impregnated with a quite novel buffer solution  
comprising an ingredient having a high buffering performance. Thus, an ad-  
vantage occurs in that a suitable quantity of pure water can be added in the  
35 calibration.

In addition, the sheet for use in the calibration of the ion-concentration  
and the like measuring device and the method of calibrating the ion-con-  
centration and the like measuring device according to the present invention  
can be applied not only to the above described pH meter but also to general  
40 instruments, in which various kinds of ion electrodes, such as a pNA-sensor.

- 1    pK-sensor, pCl-sensor, pCa-sensor and biosensor, are required to be calibrated by the use of a calibrating solution.

- Besides, the electrode to be calibrated is not limited to the above described flat surface. The shapes of the conventional glass electrode, reference electrode,  
5    composite electrode and various kinds of ion electrodes may be used. And, the present invention can be applied to the simplified calibration of the electrodes of these ion-concentration measuring devices.

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1 CLAIMS:

1 A method for measuring the ion concentration and the like of a sample,  
characterized by impregnating a liquid-absorbing material with the sample  
and bringing said impregnated material into contact with the measuring  
5 portion of a device for measuring the ion concentration and the like.

2 The method according to claim 1, characterized in that said liquid-  
absorbing material is directly impregnated with a liquid sample.

10 3. The method according to claim 1, characterized in that said liquid-  
absorbing material is pre-impregnated with a liquid which does not influence  
the measurement and brought into contact with a comparatively dry sample.

15 4. The method according to claim 1, characterized in that said liquid-  
absorbing material is pre-impregnated with a liquid which does not influence  
the measurement and brought into contact with a gaseous sample.

5 The method according to claim 1, characterized in that a powdery or solid  
sample is placed on the upper surface of the liquid-absorbing material which is  
20 brought into contact with the measuring unit, and a liquid which does not  
influence the measurement is added dropwise to the powdery or solid sample.

6 The method according to claims 3 to 5, characterized in that the pre-  
impregnating liquid which does not influence the measurement is distilled  
25 water or a 0.1% NaOH solution which is adjusted to a pH to about 5.0 with HCl.

7. The method according to claim 1, characterized in that the measurement  
is carried out by means of a device for measuring the ion concentration, the  
quantity of dissolved oxygen or the conductivity.  
30

8. A sheet for the calibration of the device for measuring the ion concen-  
tration and the like for use in the method according to claim 1 to 7, character-  
ized by a chemically stabilized hygroscopic water-absorbing sheet eluting no  
ion impregnated with a calibrating buffer solution and thereafter dried.  
35

9. A method for calibrating the device for measuring the ion concentration  
and the like used according to claims 1 to 7, characterized in that the sheet  
according to claim 8 is brought into contact with an electrode of the device for  
measuring the ion concentration and the like and a predetermined quantity of  
40 pure water is added dropwise to said sheet.

Fig. 1

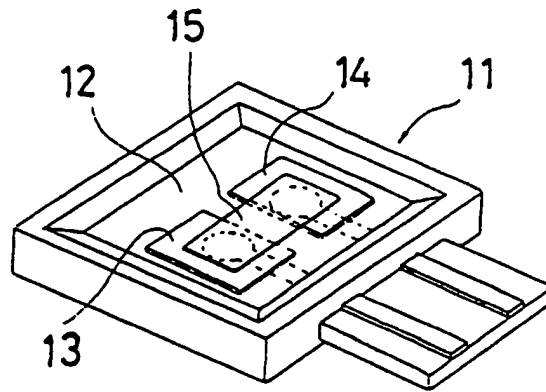


Fig. 2

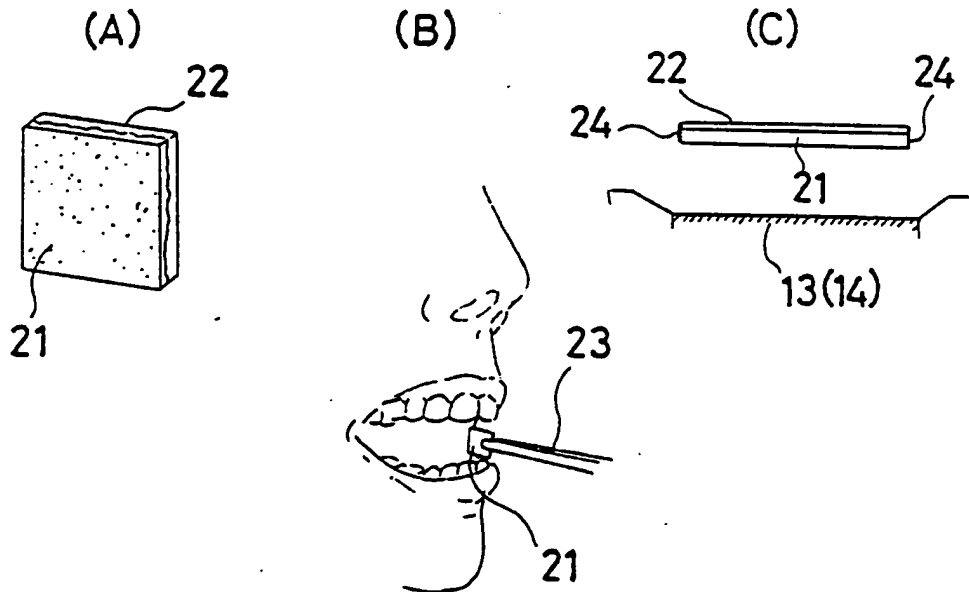


Fig. 3

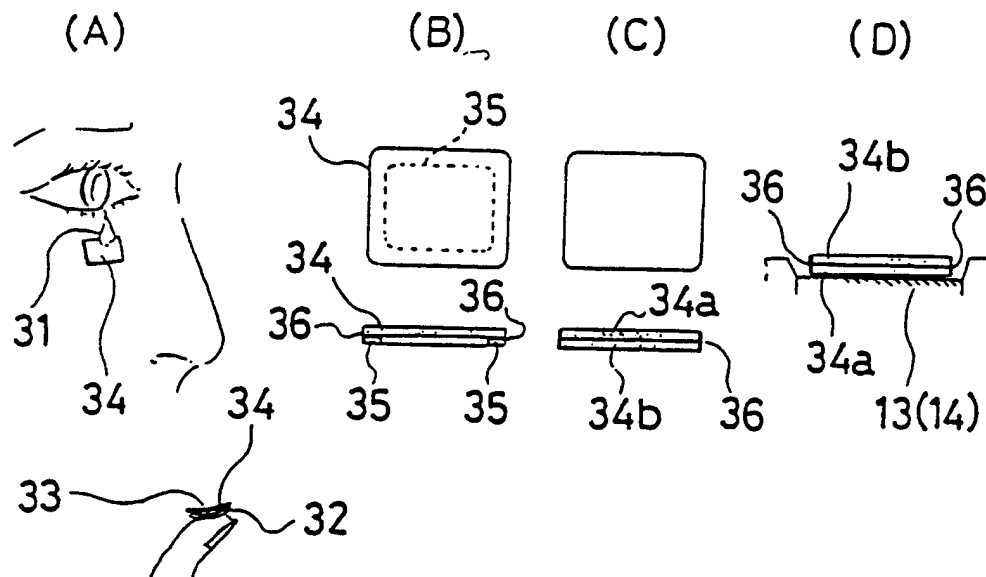


Fig. 4

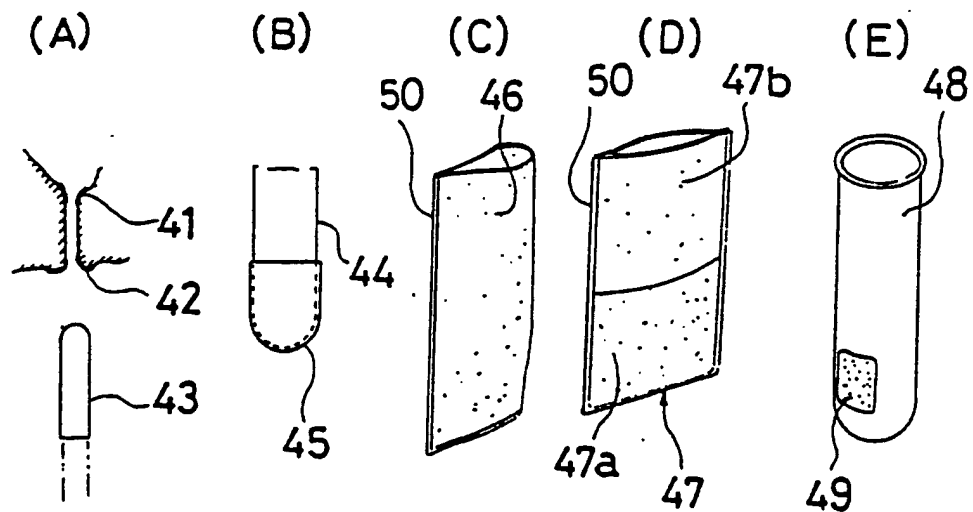




Fig.5

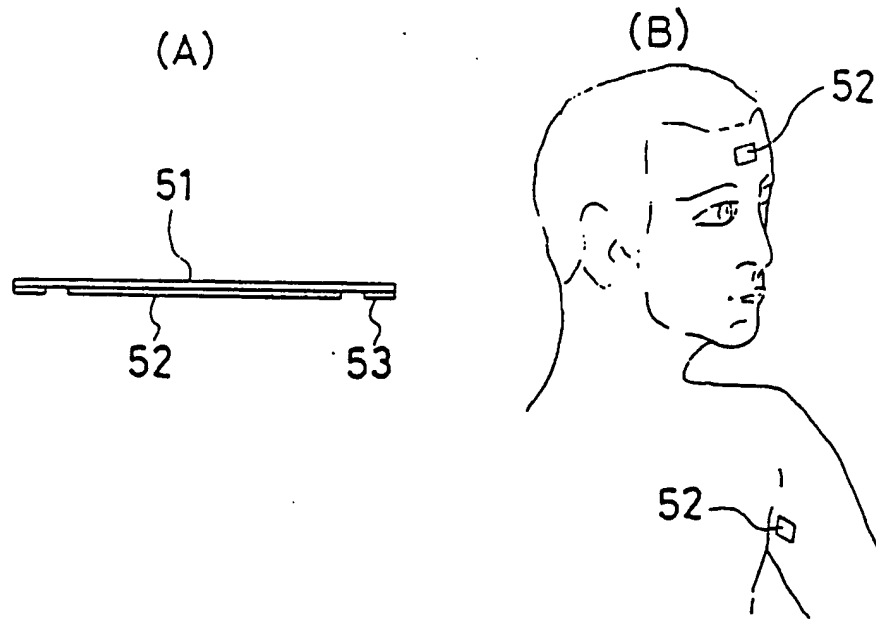


Fig.6

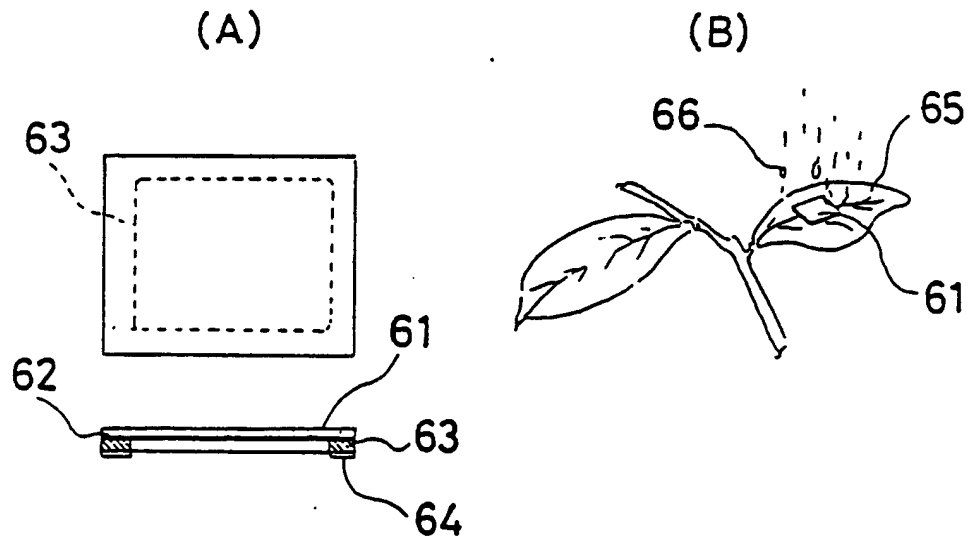


Fig. 7

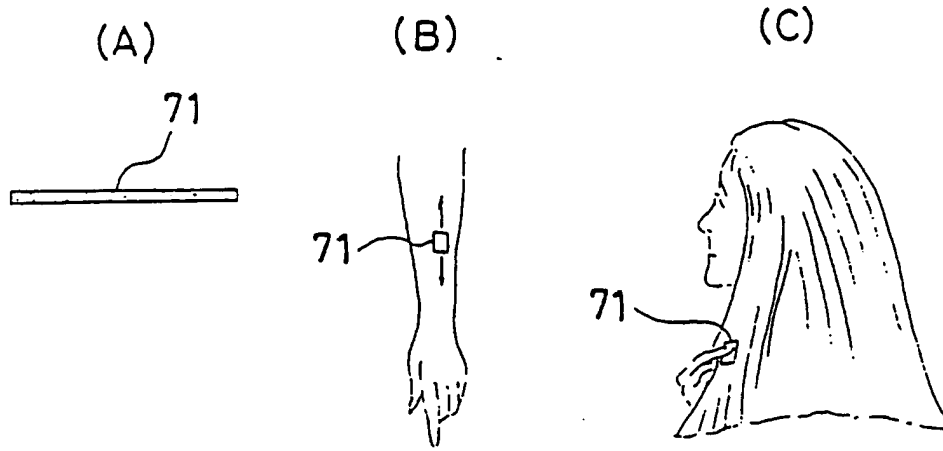


Fig. 8

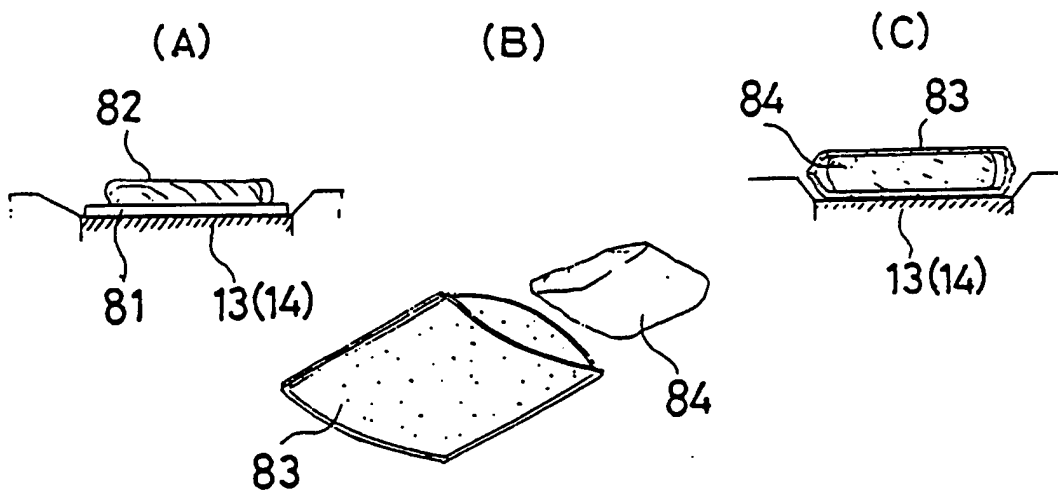


Fig. 9

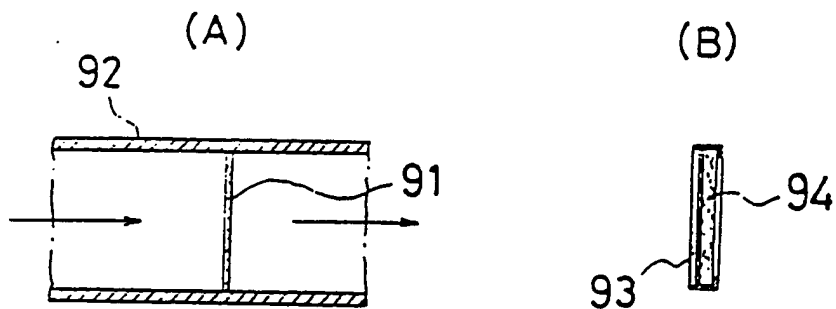
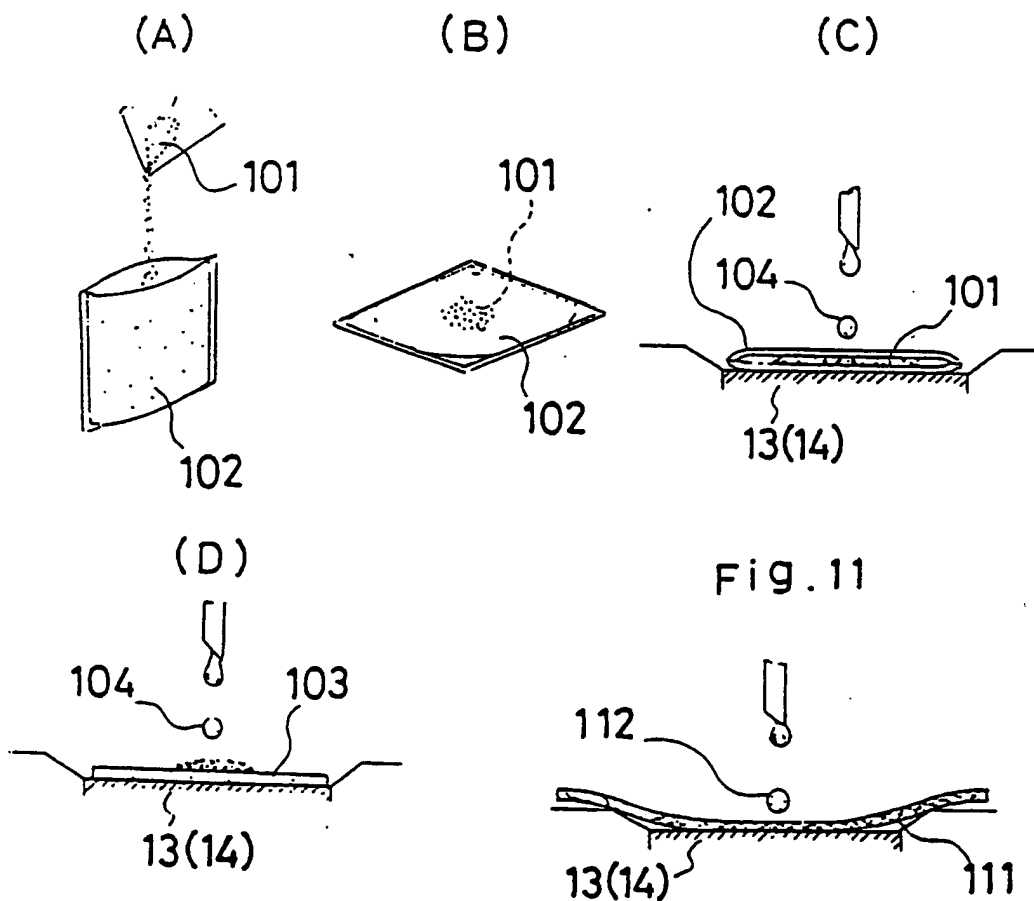
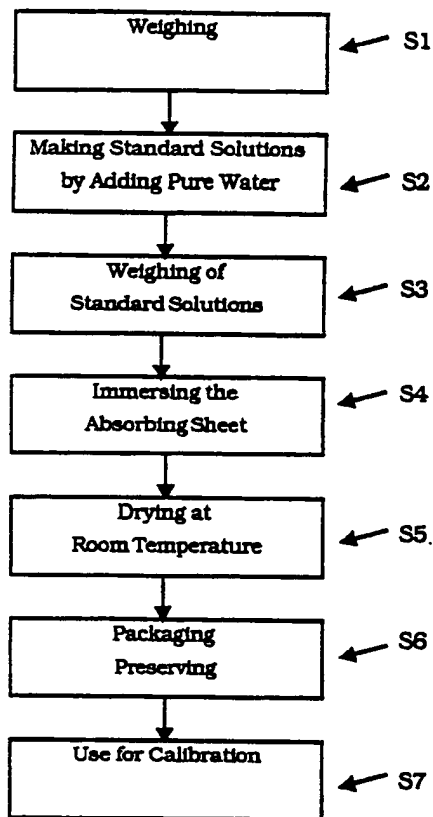


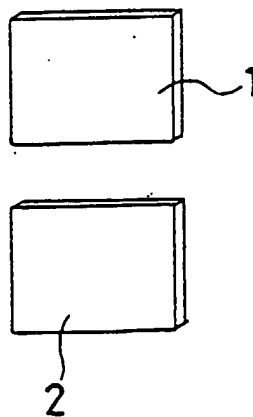
Fig. 10



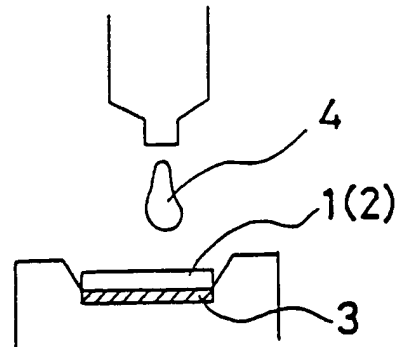
**Fig.12**



**Fig.13**



**Fig.14**



**Fig.15**  
**PRIOR ART**

